Acetate Titration

1.) Accurately weigh out about 0.7 g sample and place sample in an Erlenmeyer flask.

2.) Dissolve in H$_2$O.

3.) Titrate with 1 M HCl (use the standardized HCl from the stock room.), using erythrosin as the indicator.

4.) Calculate Experimental Molar Mass and compare to the molar mass of unknown

   a. Calculate moles = ($M_1$)$V_1$  
      
      $M_1$ = molarity of the HCl  
      $V_1$ = volume (in liters) used in the titration

   b. moles = initial weight used/ Experimental Formula Weight

      Moles are known from “a”

Compare Experimental Formula Weight to known Formula Weight of sample.